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#Jenny



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Cool! I'am really happy

#Markus Jensen



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#Diego Butler



so many fake sites. this is the first one which worked! Many thanks

Step 1 of 4
(a)
Consider the following data:
Amount of electrons, $n = 6.48 \times 10^{17}$
Calculate the number of coulombs represented by 6.48×10^{17} electrons.
 $Q = (6.48 \times 10^{17})(1.6 \times 10^{-19})$
 $= 10.37 \times 10^{-2}$
 $= 0.1037 \text{ C}$
Therefore, the number of coulombs, Q represented by 6.48×10^{17} electrons is **0.1037 C**.

Step 2 of 4
(b)
Consider the following data:
Amount of electrons, $n = 1.24 \times 10^{18}$
Calculate the number of coulombs represented by 1.24×10^{18} electrons.
 $Q = (1.24 \times 10^{18})(1.6 \times 10^{-19})$
 $= 1.984 \times 10^{-1}$
 $= 0.1984 \text{ C}$
Therefore, the number of coulombs, Q represented by 1.24×10^{18} electrons is **0.1984 C**.

Step 3 of 4
(c)
Consider the following data:
Amount of electrons, $n = 2.46 \times 10^{18}$
Calculate the number of coulombs represented by 2.46×10^{18} electrons.
 $Q = (2.46 \times 10^{18})(1.6 \times 10^{-19})$
 $= 3.936 \text{ C}$
Therefore, the number of coulombs, Q represented by 2.46×10^{18} electrons is **3.936 C**.

Step 4 of 4
(d)
Consider the following data:
Amount of electrons, $n = 1.628 \times 10^{18}$
Calculate the number of coulombs represented by 1.628×10^{18} electrons.
 $Q = (1.628 \times 10^{18})(1.6 \times 10^{-19})$
 $= 2.6048 \times 10^{-1}$
 $= 26.048 \text{ C}$
Therefore, the number of coulombs, Q represented by 1.628×10^{18} electrons is **26.048 C**.

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