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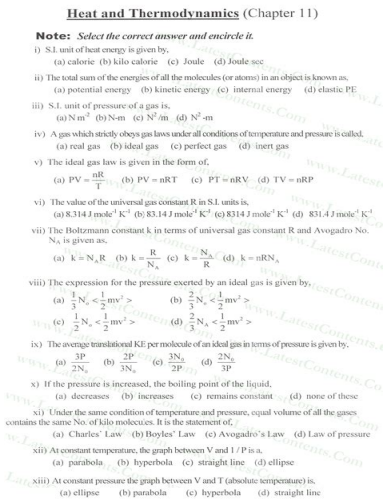
so many fake sites. this is the first one which worked! Many thanks



Heat and Thermodynamics (Chapter 11)

Note: *Select the correct answer and encircle it.*

i) S.I. unit of heat energy is given by,
(a) calorie (b) kilo calorie (c) Joule (d) Joule/sec

ii) The total sum of the energies of all the molecules (or atoms) in an object is known as,
(a) potential energy (b) kinetic energy (c) internal energy (d) elastic PE

iii) S.I. unit of pressure of a gas is,
(a) N m^{-2} (b) N-m (c) N^2/m (d) N^2 -m

iv) A gas which strictly obeys gas laws under all conditions of temperature and pressure is called,
(a) real gas (b) ideal gas (c) perfect gas (d) inert gas

v) The ideal gas law is given in the form of,
(a) $PV = \frac{nR}{T}$ (b) $PV = nRT$ (c) $PT = nRV$ (d) $TV = nRP$

vi) The value of the universal gas constant R in S.I. units is,
(a) 8.314 J mole^{-1} K^{-1} (b) 83.14 J mole^{-1} K^{-1} (c) 8314 J mole^{-1} K^{-1} (d) 831.4 J mole^{-1} K^{-1}

vii) The Boltzmann constant k in terms of universal gas constant R and Avogadro No. N_A is given as,
(a) $k = N_A R$ (b) $k = \frac{R}{N_A}$ (c) $k = \frac{N_A}{R}$ (d) $k = nRN_A$

viii) The expression for the pressure exerted by an ideal gas is given by,
(a) $\frac{1}{3}N_v < \frac{1}{2}mv^2 >$ (b) $\frac{2}{3}N_v < \frac{1}{2}mv^2 >$
(c) $\frac{1}{2}N_v < \frac{1}{2}mv^2 >$ (d) $\frac{2}{3}N_v < \frac{1}{2}mv^2 >$

ix) The average translational KE per molecule of an ideal gas in terms of pressure is given by,
(a) $\frac{3P}{2N_0}$ (b) $\frac{2P}{3N_0}$ (c) $\frac{3N_0}{2P}$ (d) $\frac{2N_0}{3P}$

x) If the pressure is increased, the boiling point of the liquid,
(a) decreases (b) increases (c) remains constant (d) none of these

xi) Under the same condition of temperature and pressure, equal volume of all the gases contains the same No. of kilo molecules. It is the statement of,
(a) Charles' Law (b) Boyles' Law (c) Avogadro's Law (d) Law of pressure

xii) At constant temperature, the graph between V and 1 / P is a,
(a) parabola (b) hyperbola (c) straight line (d) ellipse

xiii) At constant pressure the graph between V and T (absolute temperature) is,
(a) ellipse (b) parabola (c) hyperbola (d) straight line

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